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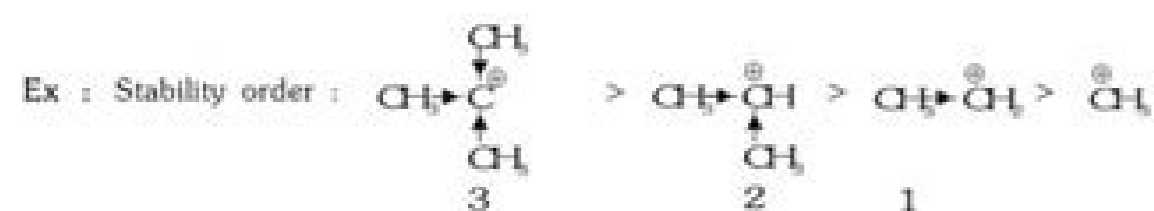
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(c) Therefore

$$\text{Energy} \propto \text{charge} \propto \frac{1}{\text{stability}}$$

$$\text{Stability of Carbocation} \propto \text{No. of +I groups}$$

$$\text{Stability of carbocation} \propto \frac{1}{\text{No. of -I groups}}$$



Sol. Reason : More no. of +I group, more stable carbocation.

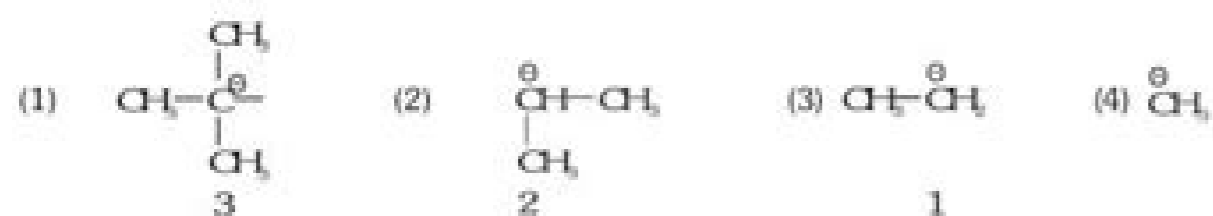
(2) Stability of carbanion :

- (a) If number of - I groups increases then stability of carbanion increases.  
 (b) If number of +I groups decreases then stability of carbanion increases.  
 (c) Therefore

$$\text{Stability of Carbanion} \propto \text{No. of -I groups}$$

$$\text{Stability of Carbanion} \propto \frac{1}{\text{No. of +I groups}}$$

Example :



More No. of +I group, Less stable carbanion.

So stability order  $1 < 2 < 3 < 4$

Example :



Minimum distance of -F.

Maximum -I of -F.

Minimum negative charge.

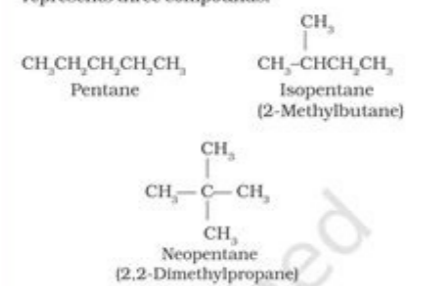
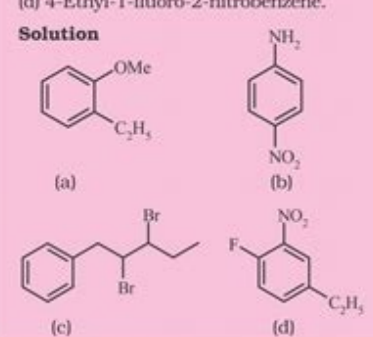
Maximum stable.

So stability order  $1 > 2 > 3$

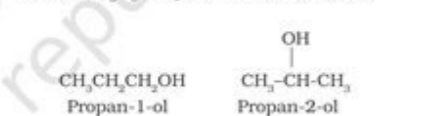
1. A series of compounds is shown in the following diagram. The compounds are arranged in order of increasing boiling point. The boiling point of the first compound is 36°C and that of the last compound is 174°C. Identify the compounds A, B, C, D, E, F, G, H, I, J, K, L, M, N, O, P, Q, R, S, T, U, V, W, X, Y, Z.

**Problem 12.10**

Write the structural formula of:  
 (a) o-Ethylaniline, (b) p-Nitroaniline,  
 (c) 2,3-Dibromo-1-phenylpentane,  
 (d) 4-Ethyl-1-bromo-2-nitrobenzene.



(ii) **Position Isomerism:** When two or more compounds differ in the position of substituent atom or functional group on the carbon skeleton, they are called position isomers and this phenomenon is termed as position isomerism. For example, the molecular formula  $\text{C}_3\text{H}_7\text{O}$  represents two alcohols:



(iii) **Functional group isomerism:** Two or more compounds having the same molecular formula but different functional groups are called functional isomers and this phenomenon is termed as functional group isomerism. For example, the molecular formula  $\text{C}_3\text{H}_6\text{O}$  represents an aldehyde and a ketone:

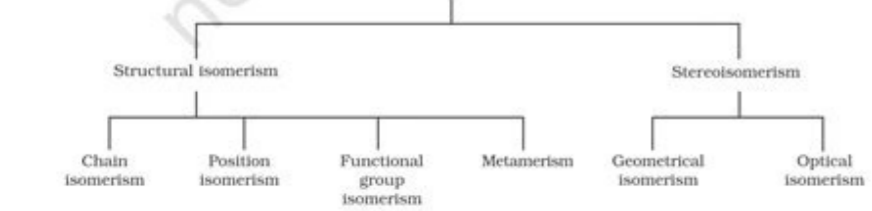
**12.6 ISOMERISM**

The phenomenon of existence of two or more compounds possessing the same molecular formula but different properties is known as isomerism. Such compounds are called as isomers. The following flow chart shows different types of isomerism.

**12.6.1 Structural isomerism**

Compounds having the same molecular formula but different structures (in terms of which atoms are linked) are classified as structural isomers. Some typical examples of different types of structural isomerism are given below:

(i) **Chain isomerism:** When two or more compounds have similar molecular formula but







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